| S.# | Question | Answer Option (A) | Answer Option (B) | Answer Option (C) | Answer Option (D) |
|---|---|---|---|---|---|
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| | | | | | |
| 1 | The mass of one mole of electron is | 1. 8 mg | .184mg | .54 mg | .64 mg |
| 2 | In the gaseous state, the distance between the molecules is ___ times _____ than their diameters. | 2 ,lesser | 3 ,lesser | 3 ,greater | 2 ,greater |
| 3 | The value of Planck's constant 'h' is | 6.626 x 1 -3 KJs | 6.262 x 1 -3 Js | 6.626 x 1 -3 Js | 6.262 x1 -3K Js |
| 4 | The radius of an electron orbit in a hydrogen atom is of the order of | 1 -8 m | 1 -1 m | 1 -11 m | 1 -13 m |
| 5 | line spectrum of sodium contains ____ colored lines separated by a definite distance | one yellow | two yellow | two brown | two golden |
| 6 | The unit_ is commonly used by meteorologist | Bar | centibar | millibar | kilobar |
| 7 | Boyle's law is stated as:"The product of pressure and volume of a fixed amount gas at constant temperature is a _____ quantity | variable | constant | discreet | decreasing |
| 8 | Number of molecules in one dm3 of water will be almost | 18x6. 2x1 23 | 55.6x6. 2x1 23 | 18x6. 2x1 23 | 2x6. 2x1 23 |
| 9 | Following are example of Intramolecular forces except? | Ionic bond | Covalent bond | Metallic bond | Dipole Dipole forces |
| 10 | The properties of liquid crystals are intermediate b/w crystals and | anisotropic liquids | nematic liquids | isotropic liquids | liquids |
| 11 | Which of the following shows H-bonding? | $CH_3CH_2OH$ | $CH_3$-O-$CH_3$ | $CH_3CH_2Cl$ | All of these |
| 12 | Atoms in solids are | Loosely Packed | In random motion | Excited | Closely packed |
| 13 | The temperature at which one crystalline form changes to other is called _temperature | critical | absolute | transition | none of these |
| 14 | Which statement is 1? | Amorphous solids have sharp melting Points | Amorphous solids have zero percent arranged | Amorphous Solids have a range of boiling points | Amorphous solids have definite heat of fusion |
| 15 | Activated complex is formed due to | Pressure | Temperature | Effective collision | Ineffective collision |
| 16 | Le Chatier Principle is about | Reaction Mixture | Reactants | Equilibrium Mixture | Products |
| 17 | Reaction of Sodium with water is an example of | Reversible reaction | Endothermic | Irreversible | Slow |
| 18 | The change in concentration of reactants or products per unit time is | rate constant | reaction speed | Rate of a reaction | All of these |
| 19 | Specific Rate constant is also known as | Time Constant | relative rate constant | Instantaneous rate constant | velocity constant |
| 20 | Hydrolysis means reaction with | Oxygen | Hydrogen | Water | Air |
| 21 | Exothermic Reactions heat is | Taken in | Give out to surroundings | Neither given nor lost | None of these |
| 22 | Anything under test or observation in laboratory is called | Surrounding | System | Confined space | None of these |
| 23 | Lattice Energy of ionic compounds are found by | Hess law | Bon Habers Cycle | Law of Mass action | Law of conservation of mass |
| 24 | In Balancing Redox equation the first thing is to | balance out all the Reactants | Write the skeleton Equation | Calculate the oxidation Number | Identify the elements |
| 25 | Current in electrolysis is carried through | Free electrons | Positive Ions | Negative Ions | Both B and C |
| 26 | The hydrogen gas bubbling into one molar solution of HCl has a pressure of | 2atm | 780 mmHg | 19 Psi | 1 atm |
| 27 | The shielding effect _____ from left to right in a period | decreases | increases | remains same | no change |
| 28 | When two atoms bonded to form a stable compound, the attractive forces will __repulsive forces and the potential energy of the system will be ___- | recessive, higher | dominate, minimum | dominate,maximum | recessive, lower |
| 29 | The ionic radius is always _____than the atomic radius from which t is derived | higher | larger | moderate | smaller |
| 30 | Which one is natron? | Na2CO3 | Na2CO3.1 H2O | Na2CO3.H2O | NaHCO3 |
| 31 | The Oxidation state of Alkali metals is? | 1 | 2 | 3 | zero |
| 32 | Pure metal | Corrode slowly | Does not corrode easily | Corrode rapidly | None of these |
| 33 | Which of the methods to prevent corrosion? | Alloying two metals | Coating the surface with less reactive metal | Painting the surface with a dye | All |
| 34 | d-block elements are present _____ in the periodic table ? | Right of the periodic table | Left of the periodic table | Bottom of periodic table | Between s and p block elements |
| 35 | Who rejected vital force theory | Wohler | Berzelius | Wallus | Lyll |
| 36 | Which one of the following is not a heterocyclic compound | thiophene | anthracene | furan | pyrrole |
| 37 | structure of cyclohexane is | trigonal | pentagonal | hexagonal | none |
| 38 | Acetylene polymerizes to make | cyclohexane | benzene | hexane | toluene |
| 39 | position of double bond is called in alkenes can be identified by | bromine water | ozonolysis | baeyer's reagent | wurtz synthesis |
| 40 | alkene is | electrophile | nuleophile | both | none |
| 41 | Electrophile is _____ | Electron loving | Electron donor | Electron repellant | Electron giver |
| 42 | 2 Chloro 2 methylpropane is an example of | Primary alkyl halide | Secondary alkyl halide | Tertiary alkyl halide | None of the above |
| 43 | What is the correct order of reactivity of alkyl halides ? | R-Cl>R-Br>R-F>R-I | R-I>R-Br>R-Cl>R-F | R-I>R-Cl>R-Br>R-F | None of these |
| 44 | The classification of monohydric alcohols is due to | hydrogen atom | carbon atom | hydroxyl group | carbonyle group |
| 45 | Which of following has higher boiling and melting point | acetone | 2 - butanol | propane | 2 - methylpropene |
| 46 | The alkenes can be formed by alcohols in the presence of | acidified KMnO4 | acidified K2Cr2O7 | acidified cucl2 | pyridine |
| 47 | What is the nature of Carbon present in Aldehyde is? | Nucleophilic | Electrophilic | Neutral | all of these |
| 48 | Which of the following is the correct IUPAC name of acetone? | Propanone | 3-Propanone | 2-Butanone | 2-Propanone |
| 49 | Which of the following reaction takes place when acetone reacts with HCN ? | Electrophilic addition | Nucleophilic elimination | Nucleophilic addition | Electrophilic addition |
| 50 | Which of the following is correct general formula of aliphatic carboxylic acid ? | RCOOH | ArCOOH | PhCOOH | None of these |
| 51 | In IUPAC nomenclature, Carboxylic acids are named as_____? | Alkoxy acid | Alkanoic acid | Alkyl carboxylic acid | Alkoxylate acid |
| 52 | Carboxylic acid reduction to alcohol can be achieved by using? | $H_2$/Ni | Pd/C | $NaBH_4$ | $LiAlH_4$ |
| 53 | High molecular mass organic compounds, upon hydrolysis yield amino acids are called_____. | Carbohydrates | Lipids | Proteins | DNA |
| 54 | Most abundant protein in animals forming 25%-35% of body protein is which type of protein? | Simple proteins | Conjugated proteins | Derived proteins | None |
| 55 | A peptide having up to 10000 amino acids is called as_____? | Dipeptide | Protein | Polypeptide | Peptide |
| 56 | The temperature at which enzyme activity is maximum is called as_____? | Maximum temperature | Absolute temperature | Critical temperature | Optimum temperature |